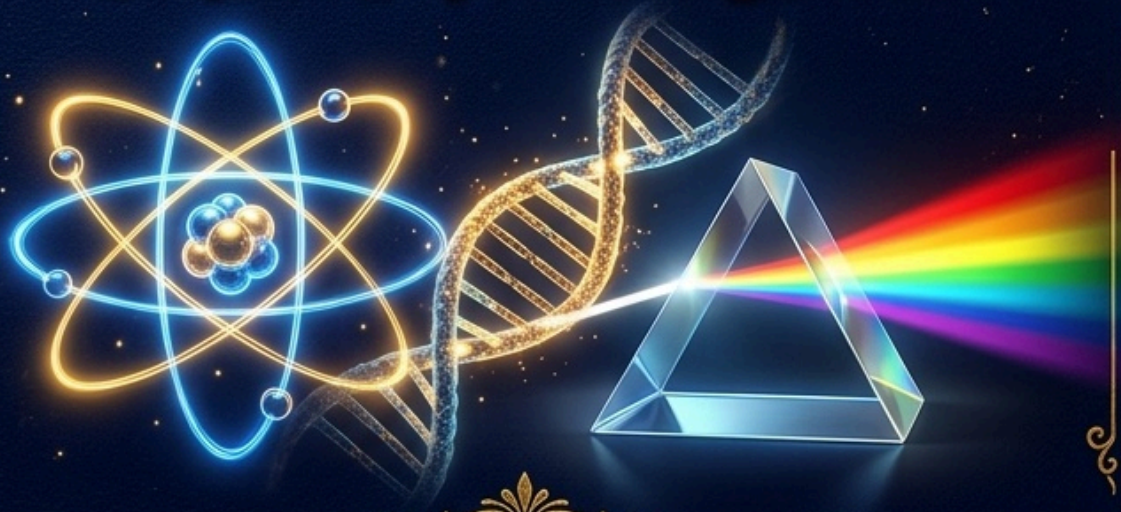
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SCIENCE HANDWRITTEN BRAHMASTRA

By Unnat Tuition Center



CBSE Class 10 Science

Complete Syllabus, Notes & Concept Maps

Chemical reaction:

The process in which new substances with new properties are formed from one or more substances is called Chemical Reaction.

Examples -

- Rusting of Iron
- Burning of wood
- Formation of Curd
- Respiration etc.

Following observations help to determine whether a chemical reaction has taken place:

1) Change in state
2) Change in colour
3) Evolution of a gas
4) Change in temperature

Reactant: The substances which take part in chemical reaction.

Product: The substances which are formed in a chemical reaction.

Chemical Equation:

- Symbolic representation of a chemical reaction is said to be **Chemical equation.**
- It involves uses of symbol of elements or chemical formula of reactant and product with mention of physical state.

- The necessary conditions such as temperature, pressure or any catalyst should be written on arrow between reactant and products.

- Example :

  Magnesium is burnt into air to form magnesium oxide can be represented as

$$Mg + O_2 \rightarrow MgO$$

Important Note

Physical state of the reactant and products are mentioned to make chemical reaction more informative.

Example we use (g) for gas, (l) for liquid, (s) for solid and (aq) for aqueous. Thus is good for chemical reactions and equations notes

Balancing Chemical Equation:

Chemical reactions and equations notes also include the mean to balance the equations by several methods. They are:

Law of Conservation of Mass -

- Mass can neither be created nor be destroyed in a chemical reaction. Thus remains conserved.

- So number of elements involved in product side will be same as reactant side.

$$H_2 + O_2 \longrightarrow H_2O$$

- In this Chemical equation number of H atoms in reactant side is 2 and that of product side is also 2. Number of O atoms in reactant side is 2 but that of product side is 1.

- To balance this equation, we have to use hit and trial method.

$$2H_2 + O_2 \longrightarrow 2H_2O$$

- Now the number of atoms in reactant side is equal to that of product side.
  Thus, balanced.

Types of Chemical Reactions

There are several types of chemical reaction based on different factors -

Based on Chemical Change:

There are total 8 types of Chemical reactions -

1) Combination Reaction – The reaction in which two or more substances combine to form a new single substance.

$$C + O_2 \longrightarrow CO_2$$

2) Decomposition reaction – The reaction in which a single substance decomposes to give two or more substances.

$$Nacl \longrightarrow Na + Cl$$

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